

# Download File PDF Explorelearning Chemical Equations Gizmo Answers Book

#Jenny



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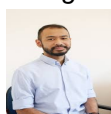
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Exploring

Name: \_\_\_\_\_ Date: \_\_\_\_\_

**Student Exploration: Chemical Equations**

**Vocabulary:** Avogadro's number, chemical equation, chemical formula, chemical reaction, coefficient, combustion, combination, conservation of matter, decomposition, double replacement, molar mass, molar volume, product, reactant, single replacement, subscript

**Prior Knowledge Questions** (Do these BEFORE using the Gizmo.)

1. A candle is placed on one pan of a balance, and an equal weight is placed on the other pan. What would happen if you lit up the candle and waited for a while? \_\_\_\_\_
2. Suppose the candle was placed in a large, sealed jar that allowed it to burn for several minutes before running out of oxygen. The candle and jar are balanced by an equal weight. In this situation, what would happen if you lit up the candle and waited? \_\_\_\_\_

**Gizmo Warm-up**

Burning is an example of a **chemical reaction**. The law of **conservation of matter** states that no atoms are created or destroyed in a chemical reaction. Therefore, a balanced **chemical equation** will show the same number of each type of atom on each side of the equation.

To set up an equation in the Chemical Equations Gizmo™, type the **chemical formula** into the left box of the Gizmo. First, type in "H<sub>2</sub>O<sub>2</sub>" in the **Reactants** box and "H<sub>2</sub>O" in the **Products** box. This represents the reaction of hydrogen and oxygen gas to form water.

| Reactants                     | → | Products         |
|-------------------------------|---|------------------|
| H <sub>2</sub> O <sub>2</sub> |   | H <sub>2</sub> O |

1. Check that the **Visual** display is chosen on each side of the Gizmo, and count the atoms:
  - A. How many hydrogen atoms are on the **Reactants** side? \_\_\_\_ **Products** side? \_\_\_\_
  - B. How many oxygen atoms are on the **Reactants** side? \_\_\_\_ **Products** side? \_\_\_\_
2. Based on what you see, is this equation properly balanced? \_\_\_\_\_

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