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The particles of suspensions are so big that the light is blocked and is not scattered. On the other hand the particles of solutions are so small that they do not scatter the rays of light.

That is why both suspensions and solutions do not show Tyndall effect. But the particles of colloids can scatter the rays of light and exhibit the Tyndall effect.

(vi) Write two/four characteristics of a colloid.
Ans. Characteristics of a Colloid:
(1) The particles are large consisting of many atoms, ions or molecules.
(2) A colloid appears to be a homogeneous mixture, hence, they are not true solutions. Particles do not settle down for a long time, therefore, colloids are quite stable.
(3) Particles are large but can't be seen with naked eye.
(4) Although particles are big but they can pass through a filter paper.
(5) Particles scatter the path of light rays thus, creating the beam of light i.e. exhibit the Tyndall effect.

(vii) Write two/four properties of a suspension.
Ans. Characteristics of a Suspension:
(1) The particles are of largest size. They are larger than 10^3 cm in diameter.
(2) Particles remain un-dissolved and form a heterogeneous mixture. Particles settle down after sometime.
(3) Particles are big enough to be seen with naked eye.
(4) Soluble particles cannot pass through filter paper.
(5) Particles are so big that light is blocked and difficult to pass.

(ix) Classify the following into true solution and colloidal solution:
blood, starch solution, glucose solution, tooth paste, copper sulphate solution, silver nitrate solution.
Ans. Real Solution/True Solution:
glucose solution, copper sulphate solution and silver nitrate solution.
Colloidal solution:
blood and starch solution.

PART II (LONG/DESCRIPTIVE QUESTIONS)

Q-5 (a) What is a saturated solution and how it is prepared? (4)
(b) How many types of solutions can be prepared depending upon the nature of solute and solvent? (3)

Ans. (a) See Chapter 6 Q.No.16
(b) See Chapter 6 Q.No.23

Q-6 (a) What do you mean by percentage concentration? How it can be expressed in terms of percentage composition by four different ways? (4)
(b) Calculate the molarity of a solution which is prepared by dissolving 28.4 g of Na_2SO_4 in 500 cm³ of solution. (3)

Ans. (a) See Chapter 6 Q.No.26
(b) See Chapter 6 Example No.2

Q-7 (a) Define solubility. Which factors affect the solubility of a solute in a solvent? (4)
(b) Compare the characteristics of a solution, colloid and a suspension. (3)

Ans. (a) See Chapter 6 Q.No.35
(b) See Chapter 6 Q.No.49

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